

Quantitative chemistry

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Timer: 15 minutes. No calculator unless a question needs one. Answers are on the next page.

Q1. Why is mass conserved in a closed system?

Q2. Relative formula mass of H_2O ($\text{H}=1$, $\text{O}=16$) is

Q3. Write the Higher Combined equation linking moles, mass and M_r .

Q4. Mass of 0.50 mol of carbon ($A_r=12$) is

Q5. How many moles in 11 g of CO_2 ($M_r=44$)?

Q6. Percentage yield =

Q7. Concentration in g/dm^3 =

Q8. Atom economy uses the M_r of what in the numerator? (Higher)

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- Q1.** Atoms are not created or destroyed
- Q2.** 18
- Q3.** $\text{moles} = \text{mass (g)} / M_r$
- Q4.** 6.0 g
- Q5.** 0.25 mol
- Q6.** $(\text{actual mass} / \text{theoretical mass}) \times 100$
- Q7.** $\text{mass (g)} / \text{volume (dm}^3\text{)}$
- Q8.** The useful product

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