

Rates of reaction

Chemistry · Foundation & Higher · 8 questions · Show working

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Timer: 15 minutes. No calculator unless a question needs one. Answers are on the next page.

Q1. Write the equation for mean rate of reaction using quantity of product.

Q2. A reaction produces 40 cm³ of gas in 20 s. Calculate the mean rate in cm³/s.

Q3. Name three factors that increase rate of reaction.

Q4. Explain, using collision theory, why a higher temperature increases rate.

Q5. What does a catalyst do to activation energy?

Q6. On a graph of volume of gas vs time, how can you tell the reaction has finished?

Q7. Why does powdering a solid increase rate?

Q8. Give one way to measure rate besides collecting gas.

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- Q1.** rate = quantity of product formed / time
- Q2.** 2 cm³/s
- Q3.** Temperature, concentration, surface area / pressure / catalyst (any three)
- Q4.** Particles have more energy; more frequent collisions; more collisions have energy e activation energy
- Q5.** Provides an alternative pathway with lower activation energy
- Q6.** The line becomes horizontal / gradient is zero
- Q7.** Larger surface area; more frequent collisions
- Q8.** Loss of mass / colour change / turbidity (cross disappearing)

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